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Periodic Table With Chemistry Formulas

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<h4>CHEMISTRY FORMULAS</h4> <h5>ATOMIC AND MOLECULAR WEIGHTS</h5> <p>A. Atomic weight: The number of protons and neutrons in an atom of an element, measured on a scale where carbon-12 is exactly 12.0000 u.</p> <p>B. Molar mass: The mass of one mole of a substance, measured in g/mol.</p> <p>C. Molecular weight: The sum of the atomic weights of the atoms in a molecule.</p>	<p>D. Density: The mass per unit volume of a substance.</p> <p>E. Specific heat: The amount of heat required to raise the temperature of one unit mass of a substance by one degree Celsius.</p>	<p>F. Heat capacity: The amount of heat required to raise the temperature of a substance by one degree Celsius.</p>	<p>G. Latent heat: The amount of heat required to change the state of a substance without changing its temperature.</p>	<h5>TEMPERATURE CONVERSION</h5> <p>A. Converting between Fahrenheit and Celsius:</p> $T_{(C)} = \frac{5}{9}(T_{(F)} - 32)$ $T_{(F)} = \frac{9}{5}T_{(C)} + 32$ <p>B. Rankine: The temperature scale where zero is absolute zero.</p> <p>C. Kelvin: The SI unit of temperature.</p>	<h5>IDEAL GAS LAWS</h5> <p>A. Boyle's law: Pressure is inversely proportional to volume at constant temperature.</p> $P_1V_1 = P_2V_2$ <p>B. Charles' law: Volume is directly proportional to temperature at constant pressure.</p> $\frac{V_1}{T_1} = \frac{V_2}{T_2}$ <p>C. Gay-Lussac's law: Pressure is directly proportional to temperature at constant volume.</p> $\frac{P_1}{T_1} = \frac{P_2}{T_2}$ <p>D. Combined gas law:</p> $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$ <p>E. Ideal gas law:</p> $PV = nRT$ <p>F. Dalton's law of partial pressures: The total pressure is the sum of the partial pressures.</p> $P_{total} = P_1 + P_2 + \dots + P_n$ <p>G. Avogadro's law: Volume is directly proportional to the number of moles at constant temperature and pressure.</p> $\frac{V_1}{n_1} = \frac{V_2}{n_2}$ <p>H. Molar volume: The volume occupied by one mole of a gas at standard temperature and pressure (STP) is 22.4 L.</p>	<h5>CHEMICAL EQUILIBRIUM</h5> <p>A. Equilibrium constant expression:</p> <p>B. Le Chatelier's principle: A system at equilibrium will respond to a stress by shifting the equilibrium to counteract the stress.</p> <p>C. Reaction quotient (Q): The ratio of product concentrations to reactant concentrations, each raised to its stoichiometric coefficient.</p> <p>D. Gibbs free energy change (ΔG): The maximum work that can be done by a system at constant temperature and pressure.</p>	<h5>ACID-BASE CHEMISTRY</h5> <p>A. pH and pOH: The negative logarithm of the concentration of hydrogen ions and hydroxide ions, respectively.</p> $pH = -\log[H^+]$ $pOH = -\log[OH^-]$ <p>B. Acid dissociation constant (K_a): The equilibrium constant for the dissociation of an acid.</p> $K_a = \frac{[H^+][A^-]}{[HA]}$ <p>C. Base dissociation constant (K_b): The equilibrium constant for the dissociation of a base.</p> $K_b = \frac{[OH^-][B^+]}{[BOH]}$ <p>D. Henderson-Hasselbalch equation:</p> $pH = pK_a + \log\left(\frac{[A^-]}{[HA]}\right)$	<h5>REDOX REACTIONS AND ELECTROCHEMISTRY</h5> <p>A. Standard reduction potential (E°): The potential for a half-cell reaction.</p> <p>B. Nernst equation: Relates the cell potential to the concentrations of the species in the half-cells.</p> $E = E^\circ - \frac{RT}{nF} \ln Q$ <p>C. Gibbs free energy change (ΔG): The maximum work that can be done by a cell.</p> $\Delta G = -nFE$ <p>D. Faraday constant (F): The charge of one mole of electrons.</p> $F = 96,485 \text{ C/mol}$
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Synopsis

SparkCharts®: The information you need—concisely, conveniently, and accurately. Created by Harvard students for students everywhere, these study companions and reference tools cover a wide range of college and graduate school subjects, from Business and Computer Programming to Medicine, Law, and Languages. They'll give you what it takes to find success in school and beyond. Outlines and summaries cover key points, while diagrams and tables make difficult concepts easier to grasp. This two-page chart is a perfect reference for homework and problem sets. On side one, the chart lists the most important chemical formulas and provides quick refreshers on significant figures and balancing equations. Side two includes a beautiful periodic table that gives the following information for each element: Name Atomic number Atomic symbol Atomic mass Oxidation states (most stable state in bold) Electronegativity Electron affinity First ionization potential Atomic radius Electron shell configuration

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Very useful for my chem class at the university of Minnesota

A nice-looking periodic table for quick reference! What more could you want?

Item arrived promptly and as described! Would use again!

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